ALQUSOUR ACADEMY

اكاديمية القصور

CHEMISTRY

103

| Subject | First Exam - Past Years & Suggested Questions |

خاص
للفصل الدراسي الأول
2013 - 2012

Prod. Date
13/10/2012

ALQUSOUR ACADEMY

Pages
7

Price
20

وردت الاخطاء التاليه في مأرة الامتحان الاول

Chapter 1

Page 11:

Q2 .2. (13.7+0.027) / 8.221

A.   1   B.2   C.3   D.4

Chapter 4

Page 7

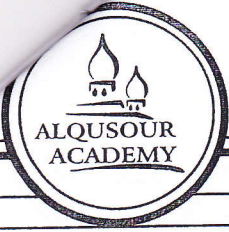**Ex 2: V=n/M ……0.0098/0.61=0.016**

Page 9

Q3.Calculate the volume of 0.5 M HCl solution needed to neutralize with 10 mL Of a 0.2 M Ba(OH)$_2$?

A. 2 ml   B. 4 ml   C. 6 ml   D. 8 ml

ملاحظة هامة : نود إعلامكم بأن مادة الامتحان لأول تكوين من 4 chapters بالإضافة الى لا سنة .

1-some molecules move with speeds approaching the "escape velocity" from earth, which is 7.0 miles per second. What is this speed in cm /h? (1 mile =1609m)

A.313   B. 4.1×10^5 C. 4.1×10^9 D. 1.1×10^6

2- Lead melt at 601.0 C. what temperature isthis in F?

A.302 F   B. 365F   C. 1,050 F   D. 1,114 F

3- The number 1.050×10^9 has how many significant figuers?

A.2   B. 3   C. 4   D. 3

4- How many significant figures are there in 1.3070×10^{2}?

A.2   B. 3   C. 4   D. 5

5- Do the indicated arithmetic and give the answer to the correct number of the significant figuers. $(1.500 + 61.326) \div 2.01 =$

A.32.0   B. 31.2   C. 31.25   D. 31.3

6- When 7.02 C is converted to the fahrenheit sacle, how many significant figures are there in the F result?

A.1   B. 2   C. 3   D. 4

7- The density of mercury is 13.6 g/cm^3. How many liter dose 251g of Hg occupy?

A.18.5L   B. 54.9L   C. 1.85×10^{-2} L   D. 5.42×10^{-2} L

8- The density of Mg is 1.74 g/cm^3 .what is the mass of magnesium block that measure 2.5 cm× 3.5cm × 1.5 cm?

A.48g   B. 23g   C. 12g   D. 7.5g

9-Element in the group 7A are known as the?

A.chalogens   B. alkali metal   C. halogens   D. Nobel gases

10- The ions Ca^{+2} and Po_4^{-3} form a salt with the formula?

A.$CaPo_4$ B.$Ca_2(Po4)_2$ C. $Ca_3(Po_4)_3$ D. $Ca_3(Po_4)_2$

11- Which of the species make an isoelecteric pair (CL^{-1}, Ca, K, S^{-2}, Fe^{+3}) ?

A. CL^{-1} and S^{-2} B. $CL^{-1}$and Ca   C. Ca and Fe^{+3} D. none of these

12-How many moles of HCL are represented by 1.0×10^{19} HCL molecules?

A. 1.7×10^{-5} B. 1.5×10^{-3}g   C. 1.0×10^{19}g   D. 36.5

13- What is the mass of 7.80×10^{18} carbon atoms?

A. 1.30×10^{-5} g   B. 6.43×10^{3} g   C. 7.80×10^{18} g   D. 1.56×10^{-4} g

14- the number of proton, neutron and electron in S^{-2} are ..., ... and ......respectively.

A. 16,32,16   B. 16,16,16   C. 16,32,18   D. 16,16,18   E. 16,16,14

15- How many sulfer atoms are there in 21.0g of Al_2S_3?

A. 8.42×10^{22} B. 2.53×10^{23} C. 2.14×10^{23} D. 6.02×10^{23}

16- How many gram of Cl_2 can be prepared from the reaction of 16.0 g of MnO_2 and 30.0 g of HCl according to the following chemical equation ?

$$MnO_2 + 4\,HCl \longrightarrow MnCl_2 + Cl_2 + 2H_2O$$

A. 0.82   B. 5.8   C. 13.0   D. 14.6

17- Calculate the mass of excess regent remaining at the end of the reaction in which 90 g of SO_2 are mixed with 100 g of O_2 ?

$$2SO_2 + O_2 \longrightarrow 2SO_3$$

A. 11.5   B. 22.5   C. 67.5   D. 77.5

18- The first step in the process for producing nitric acid is

$$4NH_3 + 5O_2 \longrightarrow 4NO + 6H_2O$$

If the reaction of 150 g of ammonia with 150 g of oxygen gas yields 87 g of nitric oxide (NO), what is the percent yield of this reaction?

A. 100 %   B. 49 %   C. 77 %   D. 33 %

19- The molarity of an aqueous solution containing 22.5g of $C_{12}H_{22}O_{11}$ in 35.5 mL of solution?

A. 0.0657   B.1.85   C. 3.52   D. 0.104

20- What volume of concentrated solution of 6.00 M NaOH must be diluted to 200 mL to make 1.50 M solution of NaOH?

A. 0.0500 mL   B.50.0 mL   C. 45.0 mL   D. 800 mL

21- The molar mass of $Mg(ClO_4)_2.6H_2O$ is

A.132 g/mol   B.231.8 g/mol   C.295.8 g/mol   D.331.3g/mol

22-which solution contains the largest number of moles of chloride ions?

A.7.50 ml of 0.500 M $FeCl_3$ B. 4.00 ml of 1.00M NaCl

C.10.0 ml of 0.500 M $BaCl_2$ D. 30.0 ml of 0.100 M $CaCl_2$

23-the point in a titration at which the indicator color change is called the

A. set point   B. Indicator point   C. standard point

D. end point   E. equivalent point

24-round the number 3.995 to two significant figures?

A. 3.9   B. 3.0   C.4   D.4.0

25- if the reaction of unknown amount of O_2 with 13.5 g of Al produced 16.2g (actual yield) of Al_2O_3 : $4Al + 3O_2 \longrightarrow 2Al_2O_3$, calculate the mass of excess reactant left at the end of reaction if the percent yield of the reaction is 79.4 %

A. 2.7   B.6.9   C.1.2   D.3.9

KEY

Question No.	Answer	Question No.	Answer
1	A	14	D
2	D	15	B
3	C	16	C
4	D	17	D
5	D	18	C
6	C	19	B
7	C	20	B
8	B	21	D
9	C	22	A
10	D	23	D
11	A	24	D
12	A	25	A
13	D		

د. ميسـة شريد : 0799851353

م. مشدى بجبرين : 0786354640

ب. آيـة العبادي : 0786214583

ب. سريا الكباريتة : 0788078762

ب. عبدالله أبوخترة : 0786491174